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- Relate number of particles and volume using Avogadro's principle. mole: an SI base unit used to measure the amount of a substance; the amount of a pure substance that contains 6.02×10^{23} representative particles
- Relate the amount of gas present to its pressure, temperature, and volume using the ideal gas law.

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Explain why beginning scuba divers are taught never to hold their breath while ascending from deep water. As a scuba diver ascends, pressure decreases. A decrease in pressure results in an increase in volume. If a diver holds his or her breath GasesGases - Weebly Ch 13 - States of Matter. 13.1 Gases.

Kinetic-molecular theory, particle size, elastic collision, temperature, diffusion, Graham's Law of Effusion (smaller molecules of gas diffuse

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faster) pressure, barometer, pascal (Pa, kPa), atmosphere (atm), torr = mmHg, Dalton 's Law of Partial Pressures Chapter Summaries - Chemistry Matter and Change 520Chapter 13 Gases. Chapter Problems521. 46. Explain why decreased volume for the gasoline-air mixture in the cylinders of a gasoline engine, increased number of gas particles, and increased temperature lead to an increase in pressure in the cylinders. Chapter 13 Gases - An Introduction to Chemistry Name Date Class 74 Chemistry: Matter and Change • Chapter 13 Study Guide for Content Mastery In your textbook, read about gas pressure. Circle the letter of the choice that best completes the statement or answers the question.13.Pressure is

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